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Density of Gas Formulas

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List of 13 Density of Gas Formulas

Density of Gas

1) Density given Relative Size of Fluctuations in Particle Density

[Open Calculator !\[\]\(339a16584d5da0f0a3ca4e9ec17bf6a1_img.jpg\)](#)

$$\text{fx } \rho_{\text{fluctuation}} = \sqrt{\frac{\left(\frac{\Delta N^2}{V_T}\right)}{[\text{BoltZ}] \cdot K_T \cdot T}}$$

$$\text{ex } 1.6E^{10}\text{kg/m}^3 = \sqrt{\frac{\left(\frac{15}{0.63\text{m}^3}\right)}{[\text{BoltZ}] \cdot 75\text{m}^2/\text{N} \cdot 85\text{K}}}$$

2) Density given Thermal Pressure Coefficient, Compressibility Factors and Cp

[Open Calculator !\[\]\(6059a5aa8b4ca7bb793408023d6c6e42_img.jpg\)](#)

$$\text{fx } \rho_{\text{TPC}} = \frac{(\Lambda^2) \cdot T}{\left(\left(\frac{1}{K_S}\right) - \left(\frac{1}{K_T}\right)\right) \cdot (C_p - [R])}$$

$$\text{ex } 0.078506\text{kg/m}^3 = \frac{\left((0.01\text{Pa/K})^2\right) \cdot 85\text{K}}{\left(\left(\frac{1}{70\text{m}^2/\text{N}}\right) - \left(\frac{1}{75\text{m}^2/\text{N}}\right)\right) \cdot (122\text{J/K}^*\text{mol} - [R])}$$



3) Density given Thermal Pressure Coefficient, Compressibility Factors and C_v

$$fx \quad \rho_{TPC} = \frac{(\Lambda^2) \cdot T}{\left(\left(\frac{1}{K_S} \right) - \left(\frac{1}{K_T} \right) \right) \cdot C_v}$$

[Open Calculator](#)

$$ex \quad 0.08665 \text{kg/m}^3 = \frac{\left((0.01 \text{Pa/K})^2 \right) \cdot 85 \text{K}}{\left(\left(\frac{1}{70 \text{m}^2/\text{N}} \right) - \left(\frac{1}{75 \text{m}^2/\text{N}} \right) \right) \cdot 103 \text{J/K}^* \text{mol}}$$

4) Density given Volumetric Coefficient of Thermal Expansion, Compressibility Factors and C_p

$$fx \quad \rho_{vC} = \frac{(\alpha^2) \cdot T}{(K_T - K_S) \cdot C_p}$$

[Open Calculator](#)

$$ex \quad 87.09016 \text{kg/m}^3 = \frac{\left((25 \text{K}^{-1})^2 \right) \cdot 85 \text{K}}{(75 \text{m}^2/\text{N} - 70 \text{m}^2/\text{N}) \cdot 122 \text{J/K}^* \text{mol}}$$

5) Density given Volumetric Coefficient of Thermal Expansion, Compressibility Factors and C_v

$$fx \quad \rho_{vC} = \frac{(\alpha^2) \cdot T}{(K_T - K_S) \cdot (C_v + [R])}$$

[Open Calculator](#)

$$ex \quad 95.45031 \text{kg/m}^3 = \frac{\left((25 \text{K}^{-1})^2 \right) \cdot 85 \text{K}}{(75 \text{m}^2/\text{N} - 70 \text{m}^2/\text{N}) \cdot (103 \text{J/K}^* \text{mol} + [R])}$$



6) Density of Gas given Average Velocity and Pressure

$$\text{fx } \rho_{AV_P} = \frac{8 \cdot P_{\text{gas}}}{\pi \cdot ((C_{\text{av}})^2)}$$

[Open Calculator !\[\]\(e78f798d4ea5c530c9db49e7d26e6b95_img.jpg\)](#)

$$\text{ex } 0.0219\text{kg/m}^3 = \frac{8 \cdot 0.215\text{Pa}}{\pi \cdot ((5\text{m/s})^2)}$$

7) Density of Gas given Average Velocity and Pressure in 2D

$$\text{fx } \rho_{AV_P} = \frac{\pi \cdot P_{\text{gas}}}{2 \cdot ((C_{\text{av}})^2)}$$

[Open Calculator !\[\]\(05be7c7a8995decd503647c99211f7c2_img.jpg\)](#)

$$\text{ex } 0.013509\text{kg/m}^3 = \frac{\pi \cdot 0.215\text{Pa}}{2 \cdot ((5\text{m/s})^2)}$$

8) Density of Gas given Most Probable Speed Pressure

$$\text{fx } \rho_{MPS} = \frac{2 \cdot P_{\text{gas}}}{(C_{\text{mp}})^2}$$

[Open Calculator !\[\]\(fe3aebe81acea8d45108cd2768939da7_img.jpg\)](#)

$$\text{ex } 0.001075\text{kg/m}^3 = \frac{2 \cdot 0.215\text{Pa}}{(20\text{m/s})^2}$$



9) Density of Gas given Most Probable Speed Pressure in 2D 

$$\text{fx } \rho_{\text{MPS}} = \frac{P_{\text{gas}}}{(C_{\text{mp}})^2}$$

[Open Calculator !\[\]\(e2376d476d06eb31946dc01a69a4403a_img.jpg\)](#)


$$\text{ex } 0.000538\text{kg/m}^3 = \frac{0.215\text{Pa}}{(20\text{m/s})^2}$$

10) Density of Gas given Root Mean Square Speed and Pressure 

$$\text{fx } \rho_{\text{RMS}_P} = \frac{3 \cdot P_{\text{gas}}}{(C_{\text{RMS}})^2}$$

[Open Calculator !\[\]\(0b5e7e25e8775f7e7e80906ada4f0021_img.jpg\)](#)

$$\text{ex } 0.00645\text{kg/m}^3 = \frac{3 \cdot 0.215\text{Pa}}{(10\text{m/s})^2}$$

11) Density of Gas given Root Mean Square Speed and Pressure in 1D 

$$\text{fx } \rho_{\text{RMS}_P} = \frac{P_{\text{gas}}}{(C_{\text{RMS}})^2}$$

[Open Calculator !\[\]\(bd3b31712ad9bab5a241210fa6925cdd_img.jpg\)](#)

$$\text{ex } 0.00215\text{kg/m}^3 = \frac{0.215\text{Pa}}{(10\text{m/s})^2}$$



12) Density of Gas given Root Mean Square Speed and Pressure in 2D

$$\text{fx } \rho_{\text{RMS}_P} = \frac{2 \cdot P_{\text{gas}}}{(C_{\text{RMS}})^2}$$

[Open Calculator !\[\]\(d3fb9f94af8b26d1c844efa9a98805b0_img.jpg\)](#)

$$\text{ex } 0.0043 \text{kg/m}^3 = \frac{2 \cdot 0.215 \text{Pa}}{(10 \text{m/s})^2}$$

13) Density of Material given Isentropic Compressibility

$$\text{fx } \rho_{\text{IC}} = \frac{1}{K_S \cdot (c^2)}$$

[Open Calculator !\[\]\(e1d6102fe77919492c04879c8450f1f5_img.jpg\)](#)

$$\text{ex } 1.2 \text{E}^{-7} \text{kg/m}^3 = \frac{1}{70 \text{m}^2/\text{N} \cdot ((343 \text{m/s})^2)}$$



Variables Used









- **c** Speed of Sound (Meter per Second)
- **C_{av}** Average Velocity of Gas (Meter per Second)
- **C_{mp}** Most Probable Velocity (Meter per Second)
- **C_p** Molar Specific Heat Capacity at Constant Pressure (Joule Per Kelvin Per Mole)
- **C_{RMS}** Root Mean Square Speed (Meter per Second)
- **C_v** Molar Specific Heat Capacity at Constant Volume (Joule Per Kelvin Per Mole)
- **K_S** Isentropic Compressibility (Square Meter per Newton)
- **K_T** Isothermal Compressibility (Square Meter per Newton)
- **P_{gas}** Pressure of Gas (Pascal)
- **T** Temperature (Kelvin)
- **V_T** Volume (Cubic Meter)
- **α** Volumetric Coefficient of Thermal Expansion (1 Per Kelvin)
- **ΔN²** Relative Size of Fluctuations
- **Λ** Thermal Pressure Coefficient (Pascal per Kelvin)
- **ρ_{AV_P}** Density of Gas given AV and P (Kilogram per Cubic Meter)
- **ρ_{fluctuation}** Density given fluctuations (Kilogram per Cubic Meter)
- **ρ_{IC}** Density given IC (Kilogram per Cubic Meter)
- **ρ_{MPS}** Density of Gas given MPS (Kilogram per Cubic Meter)
- **ρ_{RMS_P}** Density of Gas given RMS and P (Kilogram per Cubic Meter)



- ρ_{TPC} Density given TPC (Kilogram per Cubic Meter)
- ρ_{VC} Density given VC (Kilogram per Cubic Meter)



Constants, Functions, Measurements used

- **Constant:** **pi**, 3.14159265358979323846264338327950288
Archimedes' constant
- **Constant:** **[BoltZ]**, 1.38064852E-23 Joule/Kelvin
Boltzmann constant
- **Constant:** **[R]**, 8.31446261815324 Joule / Kelvin * Mole
Universal gas constant
- **Function:** **sqrt**, sqrt(Number)
Square root function
- **Measurement:** **Temperature** in Kelvin (K)
Temperature Unit Conversion 
- **Measurement:** **Volume** in Cubic Meter (m³)
Volume Unit Conversion 
- **Measurement:** **Pressure** in Pascal (Pa)
Pressure Unit Conversion 
- **Measurement:** **Speed** in Meter per Second (m/s)
Speed Unit Conversion 
- **Measurement:** **Density** in Kilogram per Cubic Meter (kg/m³)
Density Unit Conversion 
- **Measurement:** **Compressibility** in Square Meter per Newton (m²/N)
Compressibility Unit Conversion 
- **Measurement:** **Slope of Coexistence Curve** in Pascal per Kelvin (Pa/K)
Slope of Coexistence Curve Unit Conversion 
- **Measurement:** **Thermal Expansion** in 1 Per Kelvin (K⁻¹)
Thermal Expansion Unit Conversion 
- **Measurement:** **Molar Specific Heat Capacity at Constant Pressure** in Joule Per Kelvin Per Mole (J/K**mol*)




















Molar Specific Heat Capacity at Constant Pressure Unit Conversion 

- **Measurement: Molar Specific Heat Capacity at Constant Volume** in Joule Per Kelvin Per Mole ($\text{J/K}\cdot\text{mol}$)

Molar Specific Heat Capacity at Constant Volume Unit Conversion 



Check other formula lists

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- [Average Velocity of Gas Formulas](#) 
- [Average velocity of gas and Acentric factor Formulas](#) 
- [Compressibility Formulas](#) 
- [Density of Gas Formulas](#) 
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- [Molar Mass of Gas Formulas](#) 
- [Most Probable Velocity of Gas Formulas](#) 
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